

Practice Problems- Ionic Compounds

Name the following ionic compounds:

1. NH_4Br
2. Cr_2O_3
3. $\text{Co}(\text{NO}_3)_2$
4. K_2SO_4
5. $\text{Ba}(\text{OH})_2$
6. FeCl_3
7. AlF_3
8. $\text{Fe}(\text{OH})_2$
9. $\text{Cu}(\text{NO}_3)_2$
10. $\text{Ba}(\text{ClO}_4)_2$
11. Li_3PO_4
12. Hg_2S
13. $\text{Cr}_2(\text{CO}_3)_3$
14. K_2CrO_4
15. $(\text{NH}_4)_2\text{SO}_4$
16. $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$

Now go the other way- give the formulas for the following names:

17. Potassium sulfide
18. Calcium carbonate
19. Nickel (II) perchlorate
20. Magnesium sulfate
21. Silver (I) sulfide
22. Lead (II) nitrate
23. Copper (I) oxide
24. Aluminum hydroxide
25. Cesium fluoride
26. Magnesium iodide
27. Iron (III) carbonate
28. Sodium hypobromite
29. Cobalt (II) nitrate
30. Chromium (II) acetate
31. Copper (II) perchlorate
32. Calcium hydrogen carbonate

Practice Problems- Molecular Compounds

Name these binary molecular compounds:

1. SO_2
2. PCl_5
3. N_2O_3
4. SF_6
5. IF_5
6. XeO_3
7. N_2O_5
8. BF_3
9. CCl_4
10. P_4O_6
11. SiO_2
12. O_2F_2
13. XeF_6
14. AsCl_3
15. P_2O_5
16. AsBr_3

Provide formulas for the following binary molecular compounds:

17. Silicon tetrabromide
18. Disulfur dichloride
19. Dinitrogen tetroxide
20. Tetrphosphorus hexasulfide
21. Sulfur hexafluoride
22. Phosphorus tribromide
23. Carbon tetraiodide
24. Dihydrogen monoxide
25. Phosphorus triiodide
26. Iodine monobromide
27. Diboron trioxide
28. Nitrogen trichloride
29. Carbon monoxide
30. Silicon tetrachloride
31. Dinitrogen pentoxide
32. Nitrogen dioxide